



Indigo being extracted from plant leaves

Indigo: recreating Pharaoh's dye

What links your jeans, sea snails, woad plants and the Egyptian royal family? It's the dye, indigo. Learn about its fascinating history and how you can extract it at school.

Image courtesy of gitane; image source: Wikimedia Commons

By Gianluca Farusi

In ancient Egypt, only one boat had a purple-dyed sail. It belonged to the Pharaoh and was a vibrant and powerful sign to other Nile users that they should move aside to let the royal boat pass.

Even today, deep blue, purple and crimson are traditionally associated with royalty, luxury and wealth. This is because they are difficult and expensive colours to achieve using natural dyes. Until the advent of synthetic dyes 100 years ago, natural dyes (from plant, animal or mineral sources) were the only way to colour fabrics.

The vivid purple of the Pharaoh's sail was achieved using indigo – a natural dye that can be extracted from certain plants and animals. Thanks to Roman naturalist Pliny the Elder

(see box on page 46), we even know the dyeing methods used by ancient civilisations. This activity allows your students to follow in the footsteps of these early chemists, extracting indigo from the leaves of the woad plant.

Using basic lab equipment, secondary students of all ages could carry out this extraction in one or two practical sessions. Younger students (ages 11-15) could simply do the extraction without going into much detail about the chemical reaction [although note that the reviewer suggested the activity was suitable for students aged 14+]. More advanced organic chemistry students (ages 16+) could investigate the compounds and reactions more thoroughly. The experiment is also relevant to the use of science in industry and could be used as part of a project on this topic.

All about indigo

Indigo is an organic compound that is found in three differently coloured forms – indigo itself, which is blue; purple mono-bromoindigo; and red-purple di-bromoindigo (figure 1). Naturally derived dyes can contain just one of these pigments or a mixture of two or three in variable proportions, leading to a range of colours from red to blue (Cooksey, 2001). The intensity of the colour is also influenced by the way it is processed, such as whether the dyed cloth is dried in the light or shade.

In ancient Egyptian times, the best quality indigo was extracted from *Murex* sea snails, which were once common in the coastal waters of the eastern Mediterranean. Archaeological evidence from Crete suggests that indigo extraction from *Murex* snails



Modern painting of an Egyptian royal boat, with a furled imperial purple sail. Note that 'purple' is not always what we would expect today



- ✓ Chemistry
- ✓ Biology
- ✓ Organic chemistry
- ✓ Botany
- ✓ Ages 14-19*

In this and several other *Science in School* articles, Gianluca Farusi successfully blends ancient history and chemistry in an unusual mixture – in this case, an innovative and simple practical activity to isolate the dye indigo from the leaves of the woad plant (*Isatis tinctoria*). In my experience, practical activities involving intense colours are an effective way to attract the students' attention.

The additional activities and supplemental questions suggested in the article should help the students gain yet more practical skills and also stimulate intriguing questions, as well as offering sound answers.

Vladimir Petruševski,
Republic of Macedonia

*Note that the author recommended the activity for 11- to 19-year-olds

had begun by 2000 BC. By 1000 BC the Phoenicians, a civilisation centred in what is now Lebanon, were exploiting this valuable dye. Their entire economy was based on trading Tyrian purple, a violet-purple indigo dye derived from the snail *Murex brandaris* (now known as *Bolinus brandaris*). They were so famed for it that the name Phoenician is derived from the Greek word 'phoinisso', meaning 'to make red'.

Murex-derived indigo was so expensive because 12 000 snails yield just 1.4 g Tyrian purple – enough to dye a handkerchief! The compounds used to make the dyestuff are secreted by the snail's hypobranchial gland, found between the bowels and the branchial organ. The snails appear to produce these indigo precursors as a defence response and for their antimicrobial



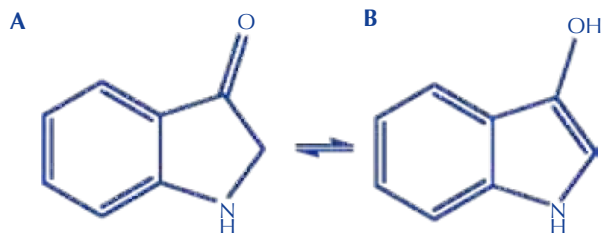
Figure 1: Structures of indigo (top), mono-bromoindigo (middle) and di-bromoindigo (bottom)

Images courtesy of Gianluca Farusi

REVIEW



Shell of the snail *Murex brandaris*, the secretions of which were first used to make indigo in 2000 BC



properties. The fresh secretions are colourless, but darken when exposed to air. According to Pliny the Elder, the dye was extracted by crushing the snails, leaving them to putrefy for three days in alkaline salt water, and then boiling them for up to ten days – imagine the smell!

The Phoenicians perfected this process over hundreds of years and discovered that by varying the species of snail, extraction methods and processing, they could produce a range of colours from red to purple to blue (table 1).

Species	Appearance	Dye precursors found in the organism	Dye name	Colour	Chemical composition of dye
<i>Bolinus brandaris</i> (formerly <i>Murex brandaris</i>)	Image courtesy of H Zell; image source: Wikimedia Commons	6-bromoindoxyl	Tyrian purple / imperial purple / argaman	Red-purple	Mainly di-bromoindigo
<i>Hexaplex trunculus</i> (formerly <i>Murex trunculus</i>)	Image courtesy of Dezidor; image source: Wikimedia Commons	Indoxyl and 6-bromoindoxyl	Royal blue / tekhelet	Blue-purple	Mixture of indigo, mono- and di-bromoindigo
<i>Indigofera tinctoria</i> (true indigo)	Image courtesy of Kurt Stüber; image source: Wikipedia Commons	Indican	Indigo	Blue	Indigo
<i>Isatis tinctoria</i> (woad)	Image courtesy of Pethan; image source: Wikipedia Commons	Isatan B	Indigo	Blue	Indigo

Table 1: Natural sources of indigo

Indigo-dyed
linen from 1873

Images courtesy of Gianluca Farusi

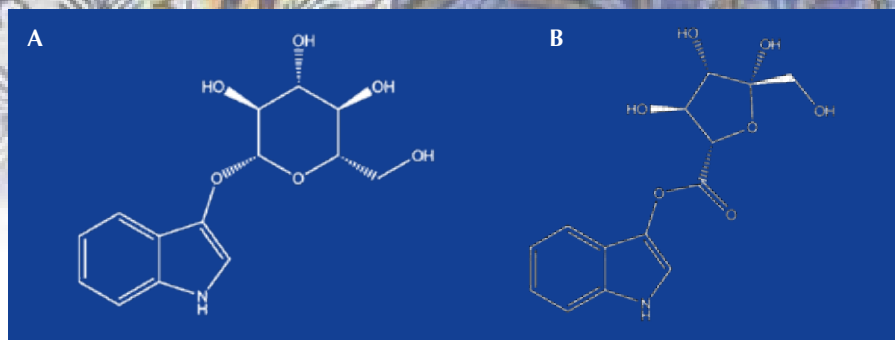


Figure 2: Indigo precursors found in plants: indican (indoxyl- β -D-glucoside, found in true indigo; A) and isatan B (indoxyl 3-ketoglucuronate, from woad; B)

Images courtesy of Gianluca Farusi

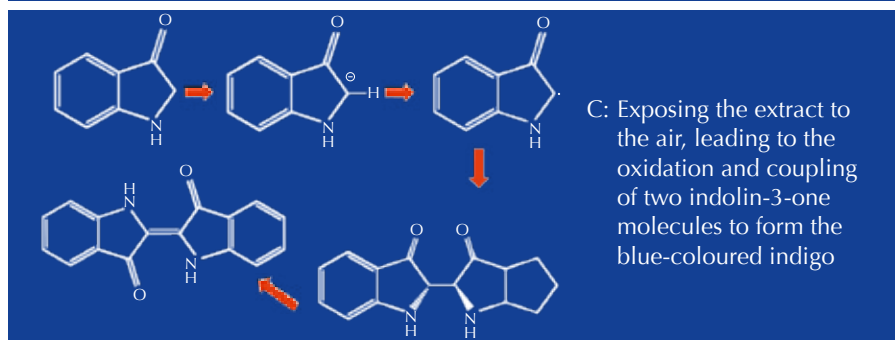
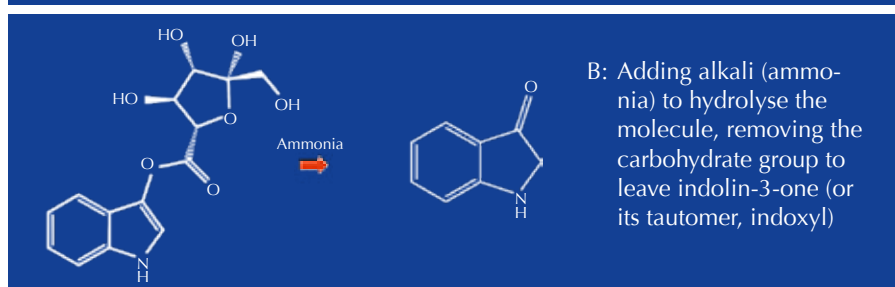


Figure 3: Chemical conversion of isatan B to indigo

Lower-quality indigo can also be extracted from certain plants, and this technique actually pre-dates the Phoenicians' snail-derived dye. In India, methods of extracting indigo from the 'true indigo' shrub *Indigofera tinctoria* were known before 2000 BC. In Europe, the dye was extracted from woad (*Isatis tinctoria*). Although these two plants contain different precursors, indican in true indigo and isatan B in woad (table 1 and figure 2), they both ultimately yield the same dye – indigo.

The production of indigo from woad takes place in three main stages:

1. Boiling the leaves to extract the water-soluble isatan B (figure 3A; steps 1-6 in the activity below).
2. Adding alkali (ammonia) to hydrolyse the molecule, removing the carbohydrate group to leave indolin-3-one or its tautomer, indoxyl (figure 3B; step 7 in the activity below).
3. Exposing the extract to the air, leading to the oxidation and coupling of two indolin-3-one molecules to form the blue-coloured indigo (figure 3C; dyeing the wool and obtaining indigo powder in the activity below).

Student activity: extracting indigo from woad leaves

Materials

For each group of students:

- Fresh woad (*Isatis tinctoria*) leaves (enough to fill a 1 l beaker)
Woad seeds are easily obtainable from garden centres, and in some countries woad is a very common plant.
- 20 wt% ammonia solution
- Three 1 l beakers
- Distilled water
- Octanoic acid (optional)
- Bunsen burner and tripod
- Stirring rod
- Sieve or strainer
- Measuring cylinders and/or 10 ml pipettes
- Heatproof gloves
- Funnel and filter paper
- Knife for cutting the leaves (optional)
- Wool for dyeing (optional)

I used a ball of white knitting wool, and some rough grey-white sheep's fleece.

Procedure

1. Fill a beaker with woad leaves. Remove the leaves from the beaker and cut or tear into small pieces.
2. Fill the beaker two-thirds full with distilled water and bring to the boil over a Bunsen burner.
3. Add a few torn leaves to the boiling water, stir and continue heating. When the water has returned to the boil, add a few more leaves. Continue adding leaves and heating until all the woad has been added.
4. When all the leaves have been added, continue boiling for 1 minute.
5. Strain all the liquid from the boiled leaves into a clean beaker.
6. Cool the extract to room temperature by floating the beaker in a shallow tray full of cold

Image courtesy of naturaldyer; image source: Flickr



Powdered indigo against a background of woad from which it was extracted

water, stirring both the extract in the beaker and the water in the tray. You now have a solution containing, among lots of other things, isatan B, which is an unstable glycoside of indoxyl.

7. When cool, measure the volume of your extract and add ammonia equivalent to 1 % of that volume, e.g. for 1 l extract, add 10 ml ammonia; for 100 ml extract, add 1 ml ammonia. The solution should immediately change from brown to brown-yellow then green. You now have a solution containing, among lots of other things, indoxyl.

To dye wool

1. Wet the wool and place in the extract for 10 minutes. Remove and leave to air dry.

To obtain indigo powder

1. Aerate the extract by pouring the solution from one beaker to another for 10 minutes.
2. Optional: if too much foam develops during aeration, add some octanoic acid, drop by drop. This acts as a surfactant.
3. Filter the extract through ordinary filter paper to obtain the indigo powder.

To use the indigo powder as a dye, it needs to be dissolved in water and mixed with a reducing agent, such as sodium hydrosulphite ($\text{Na}_2\text{S}_2\text{O}_4$). When the dyed material (e.g. wool) is exposed to air, it will turn blue.

Safety notes:

Remember that indigo and many of its precursors are dyes, so take care not to spill them on clothes or skin. When handling concentrated ammonia solution, as well as indigo and its precursors, wear gloves and chemical splash glasses, and use a fume hood.

Questions for discussion

- Why do we boil the leaves? What is happening?
- What happens when the ammonia is added?
- When you dye the wool, why does the blue colour appear as the wool dries?
- When you produce the powder, what is happening when the solution is poured from beaker to beaker? Why does the indigo precipitate?
- Can you find out how indigo is produced industrially today? What are its major uses?

Image courtesy of Marcel Douwe Dekker; image source: Flickr



Made in the 11th century, the Bayeux Tapestry celebrates the Norman conquest of England; a turning point in English history. Shown here is the Battle of Hastings at which William, Duke of Normandy, defeated Harold, Earl of Wessex, on 14 October 1066. Three plant-based dyes were used alone or in combina-

tion to produce the ten colours found in the 70 m embroidery: woad (blues and greens), madder (*Rubia tinctorum*; reds) and weld (*Reseda luteola*; yellows and dark green). To achieve the desired colour intensity, the wool was successively soaked and air-dried.



Extension activities

You could also ask your students to further investigate the chemistry of indigo and other dyes with some of the following activities.

Calculating the yield of indigo

Weigh the leaves before the extraction and the dried indigo powder. Calculate the percentage yield of indigo as follows: (mass of indigo / mass of leaves) x 100. You could compare the yield from fresh and dried leaves, or compare it with that of other indigo sources such as *Indigofera tinctoria*.

Using indigo as a dye

Our experiment involves using indigo to dye wool. Can indigo be used to dye other materials? You could compare natural materials such as cotton, linen and silk with synthetics or synthetic mixes (e.g. polyester cotton, cotton lycra).

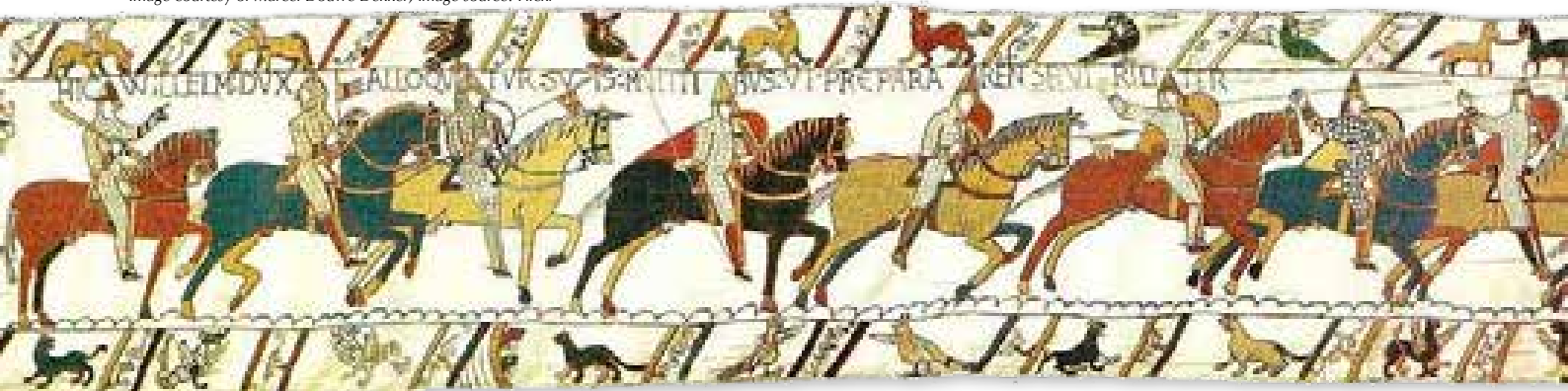
The effectiveness of indigo production

a) Which produces more indigo, larger, older leaves or younger, smaller leaves? Express the amount of indigo obtained by weight and per leaf (count leaves before extracting them). Which produces the most indigo?

- b) Grow plants with and without added nitrogen fertiliser. Aim for 2-4 g nitrogen per m² of soil (you will need to calculate the amount of fertiliser that gives this amount of nitrogen). Does adding nitrogen fertiliser increase the amount of indigo produced?
- c) Try putting the plants in the dark for a few days (cover them with black plastic). Does this have an effect on the amount of indigo produced? What can you conclude about the metabolism of isatan B in the plant?
- d) Investigate the effect of other treatments to the plants or the cut leaves.
- e) Test the effect of varying the method of the extraction, for example using different kinds of alkali.

Investigating other dyes

My students produced dyes from onion skins, from madder (Farusi, 2006), and from myrtle berries. In addition, they tested how colour-fast their dyes were by washing the dyed materials using different methods.



The Battle of Hastings begins: under the leadership of William, Duke of Normandy, the Norman forces charge the English forces, led by Harold, Earl of Wessex.



Studying chemistry with Pliny the Elder



A 12th century manuscript of *Naturalis Historia*

Image courtesy of PHICCOM; image source: Wikimedia Commons

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To learn how to use this code, see page 65.



BACKGROUND

This activity is part of a larger interdisciplinary project, developed together with 14- to 15-year-old students, to explore ancient scientific techniques. Pliny the Elder (23-79 AD) was a Roman author and naturalist. His encyclopaedia, *Naturalis Historia*, brings together much of the scientific knowledge of the time. We began each topic by discussing a passage from *Naturalis Historia* and then worked out how to recreate either the experiment described in the text or something similar. In this way, the students began in the same pre-scientific state as Pliny and, through laboratory work and discussion, gained modern scientific knowledge on each of the topics. The process motivated even the most unenthusiastic students.

Other activities in the project include recreating ancient perfumes (Farusi, 2011), preparing glass tesserae with boric acid, simulating the luminescence of the shellfish *Pholas dactylus*, and preparing iron-gall ink (Farusi, 2007).

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Resources

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